

Name _____

HONORS CHEMISTRY – 2011 SUMMER ASSIGNMENT



Welcome to Honors Chemistry!

You are about to embark on a whole new and exciting science experience. I am hoping the summer work will help with the transition you will experience in September. Chemistry is very different from Biology in that each chapter you learn will be tied into all the other topics ~ therefore; you must have a strong foundation.

Please see the summer assignment posted on our school web-site which encompasses part of the material covered in class.

A list of elements with their symbols is attached ~you are expected to have these memorized. On the second day of school you will have a quiz.

If you would like to start picking up your supplies the following is suggested: 3 in. by 5 in. cards, three ring binder (at least 1.5 inches wide) with dividers and a composition notebook for lab work. Calculators that perform logarithms and scientific notation are essential.

If you have any questions or concerns, stop by room 505 or 503. Also, our e-mails are mgillman@pthsd.k12.nj.us and criley@pthsd.k12.nj.us.

Have a great summer and looking forward to seeing you in the fall!

Mrs. Gillman and Ms. Riley

In chemistry, you will be learning the scientific names of elements and compounds, as well as completing many mathematical calculations of chemical quantities. Your summer assignment begins with learning some of these facts.

You will be quizzed on the names and symbols of the elements and polyatomic ions in this packet throughout the year (given the chemical symbol provide the properly spelled name or given the name provide the proper chemical symbol). You must know the spelling and symbol. All elements are to be written *as shown on this list* with a capital letter as the first letter and lowercase letter as the second letter. *Do not write in all caps, or in cursive.* You will also be quizzed on the metric prefixes, their meanings and the ability to convert between them

Assignment :

1. Make flash cards (on index cards not cut strips of paper) of the metric system prefixes at the bottom of this page, as well as the elements and the polyatomic ions listed on the next page.
 - a. For the metric system, put the prefix on one side and the numerical meaning on the other.
 - b. For elements and ions, put the symbol on one side and the name on the other. Please put only one element or ion per card.
 - c. *You will be given a grade for the flash cards as well as have a quiz on day 2.*
2. Complete the attached worksheet packet, to be handed in the first day of school *as a second grade*. You have been given a periodic table in case there are elements in the worksheet that are not on your list to memorize.

Metric Prefixes

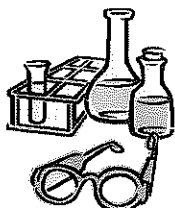
Prefix	Numerical Meaning
Kilo-(K__)	1000 base unit = 1 K__ (There are 1000 meters in a kilometer.)
BASE UNIT	The main metric unit (meter (m), liter (l) , gram (g), etc.)
deci-(d__)	10 d__ = 1 base unit (There are 10 decimeters in a meter.)
centi-(c__)	100 c__ = 1 base unit (There are 100 centimeters in a meter.)
milli- (m__)	1000 m__ = 1 base unit (There are 1000 millimeters in a meter.)
micro (μ __)	10 ⁶ μ __ = 1 base unit (There are 10 ⁶ micrometers in a meter.)
pico (p__)	10 ⁹ p__ = 1 base unit (There are 10 ⁹ picometers in a meter.)

ELEMENTS

Aluminum	Al
Argon	Ar
Barium	Ba
Beryllium	Be
Bismuth	Bi
Boron	B
Bromine	Br
Calcium	Ca
Carbon	C
Cesium	Cs
Chlorine	Cl
Chromium	Cr
Cobalt	Co
Copper	Cu
Fluorine	F
Gold	Au
Helium	He

Gallium	Ga
Germanium	Ge
Hydrogen	H
Iodine	I
Iron	Fe
Lead	Pb
Lithium	Li
Magnesium	Mg
Manganese	Mn
Mercury	Hg
Neon	Ne
Nickel	Ni
Nitrogen	N
Oxygen	O
Phosphorus	P
Platinum	Pt
Potassium	K

Radon	Rn
Rubidium	Rb
Scandium	Sc
Silicon	Si
Silver	Ag
Sodium	Na
Strontium	Sr
Sulfur	S
Titanium	Ti
Tin	Sn
Uranium	U
Xenon	Xe
Zinc	Zn

POLYATOMIC IONS

Polyatomic ions are groups of multiple atoms that have a charge (positive or negative). The symbols shown below tell you what elements are in the ion, how many atoms of each, and the charge. For example: NH_4^{+1} contains a nitrogen atom, four hydrogen atoms and the entire group has a charge of +1.

Memory Hint: If you have two ions with similar names and the only difference is the number of oxygen atoms in your ion:

-ite means smaller number of O

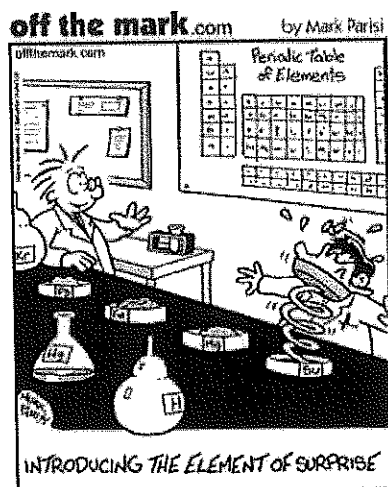
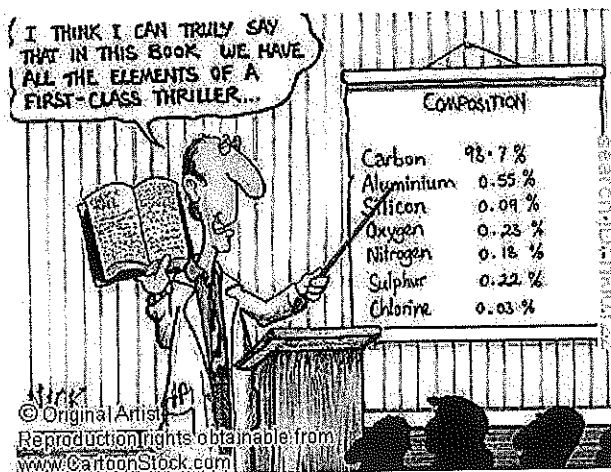
-ate means larger number of O

Hypo- (smallest) and Per- (largest) are used if there are four ions with similar names and different numbers of oxygen.

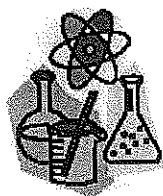
ION	NAME
NH_4^{+1}	ammonium
ClO^{1-}	hypochlorite
ClO_2^{1-}	chlorite
ClO_3^{1-}	chlorate
ClO_4^{1-}	perchlorate
CN^{1-}	cyanide
OH^{1-}	hydroxide
IO_3^{1-}	iodate
NO_3^{1-}	nitrate

ION	NAME
NO_2^{1-}	nitrite
MnO_4^{1-}	permanganate
CO_3^{2-}	carbonate
O_2^{2-}	peroxide
SO_4^{2-}	sulfate
SO_3^{2-}	sulfite
PO_4^{3-}	phosphate
$\text{CH}_3\text{COO}^{1-}$	acetate

HONORS CHEMISTRY - SUMMER ASSIGNMENT WORKSHEET PACKET



TO BE TURNED IN ON THE FIRST DAY OF CLASS
WITH YOUR FLASHCARDS.

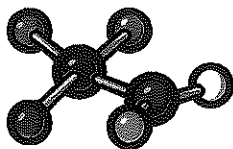


⑩ There are memorization tips and a periodic table at the back of the packet for your reference.

Memorization Tips: Elements/Symbols

Over the years, my students and I have developed several unique ways to help us remember the symbols for the elements. Be warned - some are a little out there!

Silver	Ag	If a person who is expecting a present of a gold necklace receives a silver one. He might say, " Ag , I didn't want silver!"
Gold	Au	"Hey you, I want that gold necklace!" Said with "Hey you" sounding like Au .
Bromine	Br	That brother of mine - Bro of mine!
Calcium	Ca	" Caws give milk!" Pronounced with an accent to make cows sound like it's spelled with an A.
Chlorine	Cl	"You Clean with chlorine!"
Iron	Fe	" Fe , Fi, Fo, Fum, I'm an iron man!"
Helium	He	If you breathe in helium, you will laugh! He , He, He!
Mercury	Hg	Greek mythology - Hg stands for Helmet guy !
Potassium	K	You will get Kicked out of school for the double nasty! You can't do the first three letters and cannot say the next three!
Sodium	Na	" Naw , I don't want any sodium!"
Nickel	Ni	" Nick owes me a nickel!"
Oxygen	O	" Open your mouth wide to take in oxygen!"
Lead	Pb	Pencil broke !
Silicon	Si	Silly con !
Tin	Sn	A tin roof gets hot in the Sun .
Manganese	Mn	Take first three letters - Man
Magnesium	Mg	Take first three letters - Mag



Fill in the missing symbol/name of the element. The date of discovery and the origin of the name are included for your information only. You will only be responsible for the names and symbols.

Symbol	Name	Date	Origin of Name
	aluminum	1825	Latin, alumen = astringent taste
Ar		1894	Greek, argos = neutral or inactive
	barium	1808	Greek, baryos = heavy
Bi		~1450	German, wismut = white mass
	boron	1808	Arabic, bawraq
Br		1826	Greek, bromos = stench
C		B.C.	Latin, carbo = coal
Cs		1860	Latin, caesius = blue
	chlorine	1808	Greek, chloros = green gas
Cr		1797	Greek, chroma = color
	cobalt	1735	Greek, cobolos = goblin
Cu		B.C.	Latin, cuprum
	fluorine	1886	Latin, fluere = to flow
Ga		1875	Latin name, Gaul, of France
	germanium	1886	country, Germany
Au		B.C.	Latin, aurum
He		1895	Greek, helios = the sun
H		1766	Greek, hydro genes = water former
I		1811	Greek, iodos = violet color
Fe		B.C.	Latin, ferrum
	lead	B.C.	Latin, plumbum
	magnesium	1803	Latin, magnesia = a place in Asia Minor
Mn		1774	Latin, magnes = magnet
Hg		B.C.	Latin, hydragyrum = god and planet
	neon	1898	Greek, neo = new
	nickel	1750	German, goblin
	nitrogen	1772	Latin, nitro = native soda and gen = born
O		1771	Greek, oxys = sharp and gen =

			born
P		1669	Greek, phosphoros = light bringer
	platinum	1735	Spanish, plata = silver
K		1807	Latin, kalium
	radon	1900	originates from radium
Rb		1860	Latin, rubidius = red
	scandium	1879	Scandanavian peninsula by its discoverer
	silicon	1823	Latin, silex = flint
Ag		B.C.	Latin, argentum
	sodium	1807	Latin, natrium
Sr		1808	town of Strontian, Scotland
	sulfur	B.C.	Latin, sulphur
	tin	B.C.	Latin, stannum
Ti		1791	Greek mythology, first sons of earth
U		1789	planet Uranus
Xe		1808	Greek, xenos = strange
	zinc	B.C.	German, zink = like tin

Write your answers in the blanks below

- Mg is _____
- Magnesium is _____
- Aluminum is _____
- Silicon is _____
- Fe is _____
- H is _____
- Cu is _____
- N is _____
- C is _____
- Helium is _____
- Oxygen is _____
- Copper is _____
- Calcium is _____
- Iron is _____
- Potassium is _____
- Hydrogen is _____
- Carbon is _____
- Nitrogen is _____
- O is _____
- F is _____
- Fluorine is _____
- Na is _____
- Sodium is _____

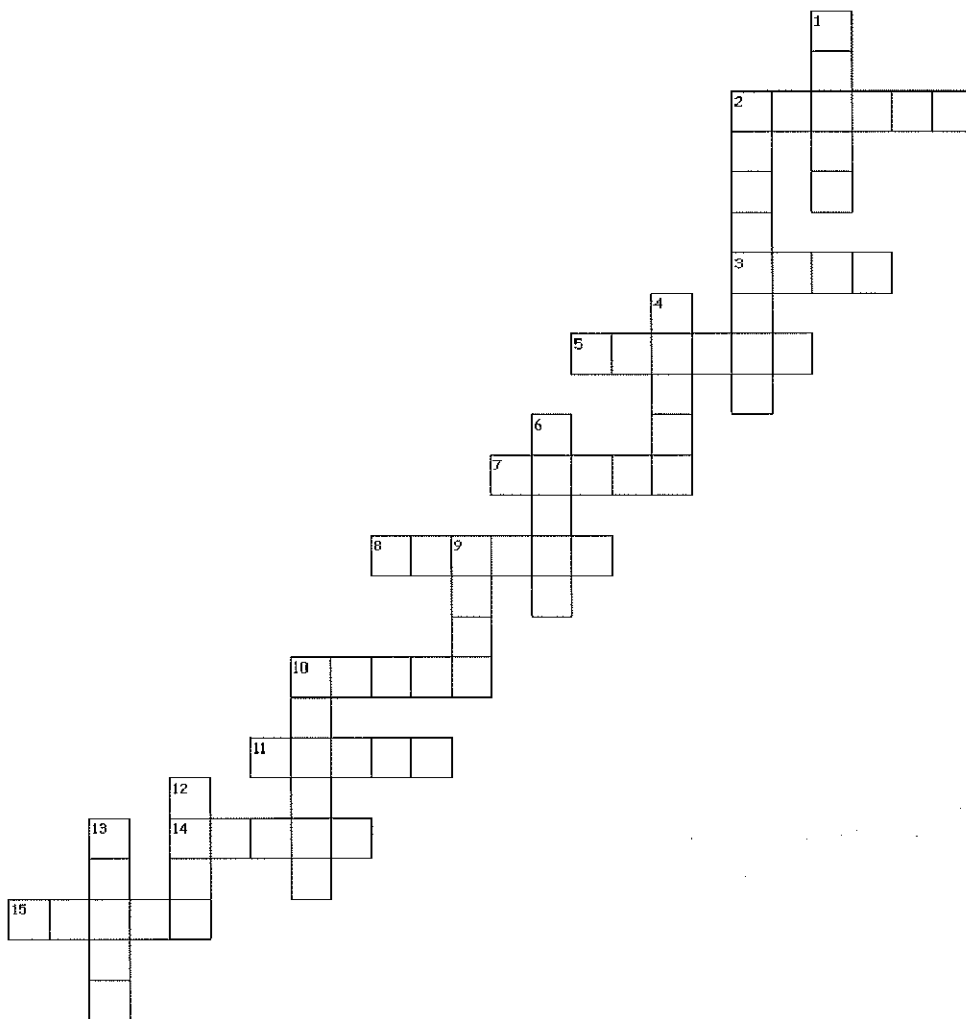
Spell the name of the following ions correctly:

1. NO_2^{1-} _____
2. CO_3^{2-} _____
3. ClO_3^{1-} _____
4. OH^{1-} _____
5. PO_4^{3-} _____
6. NH_4^{1+} _____
7. SO_4^{2-} _____
8. CN^{1-} _____
9. $\text{CH}_3\text{COO}^{1-}$ _____
10. O_2^{2-} _____
11. NO_3^{1-} _____
12. IO_3^{1-} _____
13. MnO_4^{1-} _____
14. ClO_2^{1-} _____
15. O_2^{2-} _____

Write the symbol and charge of the following ions.

1. phosphate _____
2. sulfate _____
3. cyanide _____
4. hydroxide _____
5. carbonate _____
6. nitrate _____
7. acetate _____
8. chlorate _____
9. perchlorate _____
10. hypochlorate _____
11. iodate _____
12. nitrite _____
13. sulfite _____
14. peroxide _____
15. permanganate _____

Polyatomic Ion Puzzle



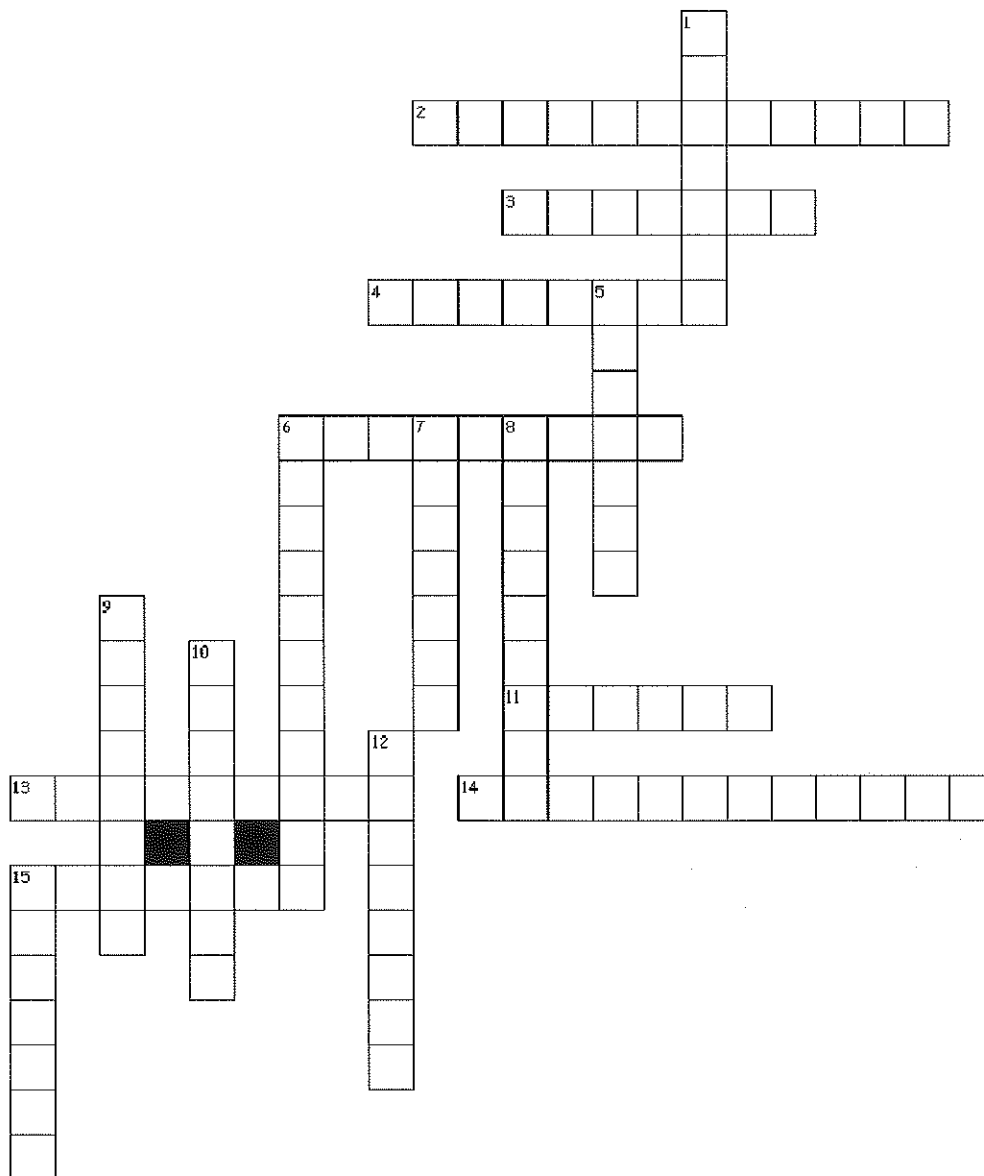
Across

- 2. chlorite
- 3. peroxide
- 5. chlorate
- 7. sulfite
- 8. permanganate
- 10. carbonate
- 11. nitrite
- 14. nitrate
- 15. sulfate

Down

- 1. hypochlorite
- 2. acetate
- 4. phosphate
- 6. iodate
- 9. hydroxide
- 10. perchlorate
- 12. cyanide
- 13. ammonium

Polyatomic Ion Puzzle



Across

- 2. ClO1-
- 3. CN1-
- 4. ClO31-
- 6. PO43-
- 11. IO31-
- 13. CO32-

- 14. MnO41-

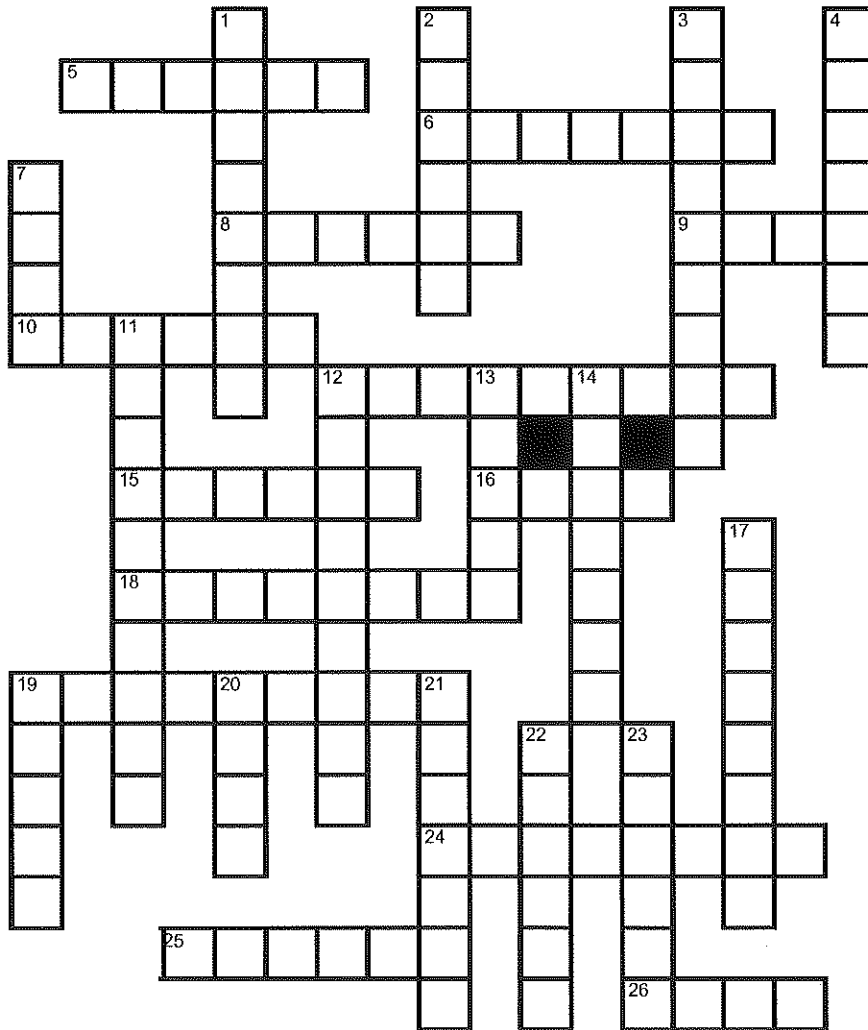
Down

- 1. SO32-
- 5. CH3COO1-
- 6. ClO41-
- 7. SO42-

- 8. OH1-

- 9. ClO21-
- 10. NH41+
- 12. O22-
- 15. NO31-

Elements of the Periodic Table

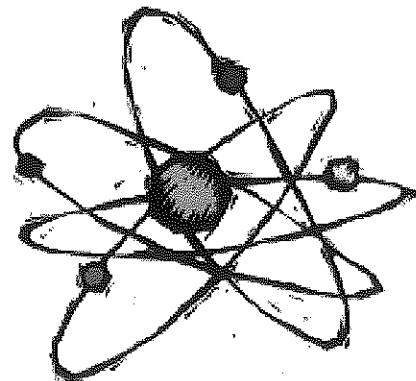


Across

- 5 He (6)
- 6 Ca (7)
- 8 O (6)
- 9 Fe (4)
- 10 Cu (6)
- 12 K (9)
- 15 Na (6)
- 16 Au (4)
- 18 H (8)
- 19 Be (9)
- 24 Cl (8)
- 25 S (7)
- 26 Ne (4)

Down

- | | |
|-----------|-----------|
| 1 N (8) | 14 Si (7) |
| 2 Ni (6) | 17 F (8) |
| 3 Al (9) | 19 B (5) |
| 4 As (7) | 20 Pb (4) |
| 7 Zn (4) | 21 Hg (7) |
| 11 P (10) | 22 Ag (6) |
| 12 Pu (9) | 23 C (6) |
| 13 Ar (5) | |



Name _____ Honors Summer Assignment

Periodic Table of the Elements Quiz

Fill in the blanks with the atomic symbols of the first 20 elements. (You do not need to add the atomic number, weight or name.) Write the element names in the blanks below.

Key:
element name
atomic number
symbol
atomic weight

		scandium 21	titanium 22	vanadium 23	chromium 24	manganese 25	iron 26	cobalt 27	nickel 28	copper 29	zinc 30	gallium 31	germanium 32	arsenic 33	selenium 34	bromine 35	krypton 36
		Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
		44.95591	47.867	50.9415	51.9961	54.93805	55.845	58.9332	58.6934	63.546	65.409	69.723	72.64	74.9216	78.96	79.904	83.798
rubidium 37	strontium 38	yttrium 39	zirconium 40	niobium 41	molybdenum 42	technetium 43	ruthenium 44	rhodium 45	palladium 46	silver 47	cadmium 48	indium 49	tin 50	antimony 51	tellurium 52	iodine 53	xenon 54
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
85.4678	87.62	88.90585	91.225	92.90638	95.94	[98]	101.07	102.9055	106.42	107.8682	112.411	114.818	118.710	121.760	127.60	126.9045	131.293

- (1) _____ (6) _____ (11) _____ (16) _____
- (2) _____ (7) _____ (12) _____ (17) _____
- (3) _____ (8) _____ (13) _____ (18) _____
- (4) _____ (9) _____ (14) _____ (19) _____
- (5) _____ (10) _____ (15) _____ (20) _____

Metric System Conversions

Fill in the blanks with the appropriate number:

- 1 liter = _____ deciliters
- 1 gram = _____ milligrams
- 1000 millimeters = _____ meters
- 1 kilogram = _____ grams
- 10^9 picometers = _____ meter
- 1 liter = _____ microliters

Metric Conversions:

You should have learned how to convert between metric prefixes in middle school. You can use ratios, or simply “move the decimal” the right number of places in the correct direction. Complete the following conversions using any method that you know.

1. How many liters are there in 145,000,000 picoliters?
2. How many grams are there in 123 kg?
3. How many centimeters are there in 921 mm?
4. What is the mass in kilograms of something that has a mass of 1926532 dg?
5. An atom has a diameter of approximately 0.10 pm. How many meters is this?
6. A 3.45 microgram sample of Uranium has a mass of how many grams?
7. What volume in kiloliters will a sample of Bay water occupy if it has a volume of 125 ml?
8. A sample of a chemical has a volume of 145 ml. How many liters is this?
9. A piece of metal has a mass of 27.9 grams.
 - a. How many kg is this?
 - b. dg?
 - c. cg?
 - d. mg?
 - e. μg ?
 - f. pg?

Name: _____

Honors Chemistry – Dimensional Analysis Worksheet

1. A football running back weighs 220 lbs; what is his weight in milligrams?
2. A piece of indium weighing 15.442 kg is placed in 49.7 mL of water in a graduated cylinder. The water level increases to 51.8 mL. What is the density of indium in g/cm^3 ?
3. The heat of combustion of ethyl alcohol is -29.8 kJ/g. What is the heat of combustion in J/kg?
4. A football field is 100 yds long, how long would the football field be in mm?
5. Vinegar contains 5.00% acetic acid by mass and has a density of 1.10 g/cm^3 . What is the mass(in grams) of acetic acid present in 10.0 L of vinegar?
6. The density of a certain substance is 4.58 g/cm^3 , what is the density in pounds per cubic foot?

7. Given the following information excerpted from the poem Jabberwocky:

There are 20 tumtum trees in the tulgey wood.

In each tulgey wood is one frumious Bandersnatch

There are 5 slithy toves in 2 borogoves

There are 2 mome raths per Jabberwocky

There are 3 Jubjub birds in 200 tumtum trees

There are 200 mome raths in each borogove

There are 5 Jubjub birds per slithy tove.

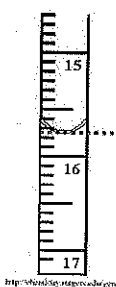
If there are 5 frumious Bandersnatches, how many Jabberwocks are there? **SHOW YOUR WORK!!!!**

2. Why was the SI unit system developed and why do we use it in Chemistry?

3. The temperature of a sample is 345K. What is this temperature in °C?

4. Indicate the measurement for the following instruments:

a.

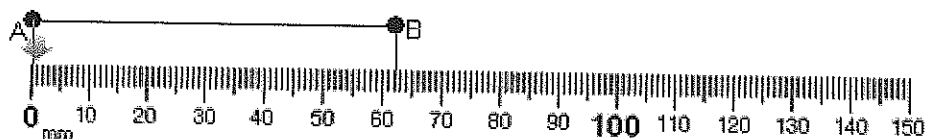


b.



**CIRCLE THE MORE
PRECISE GRADUATED
CYLINDER**

c.



Website References and Resources

Dimension Analysis

<http://www.chemtutor.com/numbr.htm#da>

Significant Figures

<http://www.chem.sc.edu/faculty/morgan/resources/sigfigs/index.html>

<http://www.ruf.rice.edu/~kekule/SignificantFigureRules1.pdf>

<http://www.wellesley.edu/Astronomy/kmcleod/Toolkit/sigfigs.html>

Please use any other resources you may find-- college sites always have great information and resources. Please email with any questions.